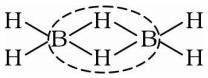
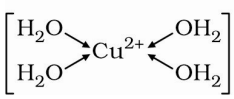
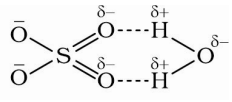
Daily Tutorial Sheet 11	Numerical Value Type for JEE Main

126.(2) 

This molecule has two (three centre two electron bonds)

127.(1) $[Cu(H_2O)_4]SO_4 \cdot H_2O$

Only one water molecule is hydrogen bonded to sulphate ion in $CuSO_4 \cdot 5H_2O$.

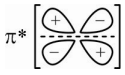



128.(3) PCl_5 has trigonal bipyramidal shape including 3-equitorial and 2-axial bonds.

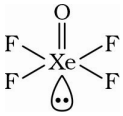
129.(2) In C_2, there are two π-bonding molecular orbitals which are completely filled.

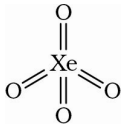
130.(4) $N_2(14) : \sigma_{1s}^2 \sigma_{1s}^{2*} \sigma_{2s}^2 \sigma_{2s}^{2*} (\pi_x^2 = \pi_y^2) \sigma_{2_{pz}}^2$

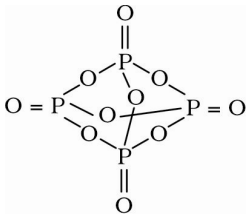
There are 4 non-bonding electrons.

131.(2) 

There are two nodal planes in it.

132.(1) 

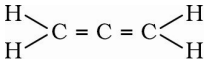
133.(4)  This molecule has four $(p\pi - d\pi)$ bonds.

134.(4)  P_4O_{10} has 4P = O bonds

135.(8) O = C = C = C = O

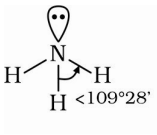
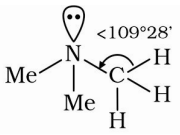
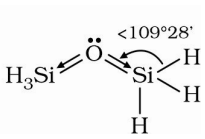
There are eight covalent bonds in C_3O_2.

136.(4) In ice, each H_2O molecule forms four hydrogen bonds.

137.(1) 

Only middle carbon atom lies in sp hybrid state.

138.(5) Species having bond angle equal to or less than 109°28' and also can act as Lewis base are :

XeO_3 AsH_3

139.(4) N_2^+, O_2, B_2, N_2^{2-} have symmetrical electronic distribution in their HOMO and are also paramagnetic.

N_2^+ paramagnetic and symmetrical electronic distribution in their HOMO.

O_2 paramagnetic and symmetrical electronic distribution in their HOMO.

B_2 paramagnetic and symmetrical electronic distribution in their HOMO.

N_2^{2-} paramagnetic and symmetrical electronic distribution in their HOMO.

O_2^{2-} diamagnetic and symmetrical electronic distribution in their HOMO.

C_2 diamagnetic and symmetrical electronic distribution in their HOMO.

C_2^{2-} diamagnetic and symmetrical electronic distribution in their HOMO.

140.(7) $B(OMe)_2$ and HCHO can not form H-bond.